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The student places a 200.0-mL sample of the solution in a 350.0mL Erlenmeyer flask and titrates the solution with a 0.1234 M solution of sodium hydroxide, where a phenolphthalein indicator is used to find the endpoint.

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A solution of liquid A is mixed with liquid B and exerts ideal behavior. At  $(103^{\circ}\text{C})$  the total vapor pressure of a solution is 413mmHg containing 1.4 moles of A and 2.0 moles of B. After you add another mole of A to the solution, the vapor pressure increases to 486mmHg. At  $(103^{\circ}\text{C})$  calculate the vapor pressure of pure A and B.

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