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SECTION 3 SHORT ANSWER

Answer the following questions in the space provided. 1. Describe the errors made by the following students in making molar solutions. a. James needs a 0.600

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M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the solid.

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representative particle diameter on the right. c solutions (a) larger than 1000 nm a suspensions (b) 1 nm to 1000 nm b colloids (c) smaller than 1 nm 2. Identify the solvent in each of the following examples:

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Answer the following questions in the space provided. 1. Solid CaCl_2 does not conduct electricity.

Explain why it is considered to be

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an electrolyte. 2. Explain the following statements at the molecular level: a. Generally, a polar liquid and a nonpolar liquid are immiscible. b.

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publications (such as books and
journal articles), critically
analyzing them, and explaining

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thorough review of solutions. Novices explain why a compound is not an electrolyte, identify types of compounds, and calculate moles, grams, and molalities in various

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MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Solid CaCl_2 does not conduct electricity.

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Section Modern Chemistry 1

Solutions CHAPTER 12 REVIEW

Solutions Teacher Notes and

Answers Chapter 12 SECTION 1

SHORT ANSWER 1. c 2. a 3. b 2.

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a. alcohol b. water c. the gels 3.
The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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Solutions - Weebly CHAPTER 12

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SECTION 2 SHORT ANSWER

Answer the following questions in the space provided. 1. The following are statements about

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the dissolving process. Explain each one at the molecular level.

a. Increasing the pressure of a solute gas above a liquid solution

increases the

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